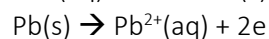
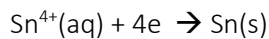
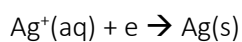
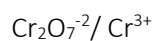


Lesson 2 – Redox. Writing oxidation and reduction half equations for more complex oxidants and reductants.

Simple oxidants such as Ag^+ are somewhat easy to write a balanced half equation for.

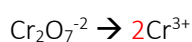


Electrons always added to the most positive side



To write a balanced half equation follow the steps

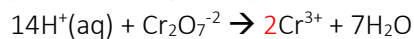
1. Balance for all elements other than O or H



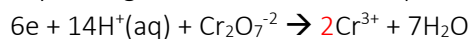
2. balance for oxygen by adding water to the side with least number of oxygens



3. balance for hydrogens by adding $\text{H}^+(\text{aq})$ to the side with least hydrogens



4. Balance for charge by adding electrons to the most positive side



1. Use the steps outlined above to give a balanced half equation, states not included, for the reaction of the the following conjugate pairs. Indicate if the reaction is an oxidation or a reduction.

Keep in mind : Reduction – electrons appear on the left

Oxidation – electrons appear on the right

